



<http://www.e-journals.net>



ISSN: 0973-4945; CODEN ECJHAO
E-Journal of Chemistry
Vol. 5, No. 2, pp.243-250, April 2008

Solar Photolysis and Photocatalytic Decolorization of Thymol Blue

FALAH H. HUSSEIN*, AHMED N. ALKHATEEB,
and JAMEEL K. ISMAIL

*College of Science, Babylon University, Hilla, Iraq.
Department of Chemistry, College of Science, Ibb University, Ibb, Yemen.
abohasan_hilla@yahoo.com

Received 1 July 2007; Accepted 20 September 2007

Abstract: The photolysis and photocatalytic decolorization of an aqueous propane-2-ol solution of thymol blue (TB) (Phenol, 4,4'-(3H-2,1-benzoxathiol-3-ylidene)bis(5-methyl-2-(1-methylethyl)-,S,S-dioxide; Thymolsulfonphthalein) ($C_{27}H_{30}O_5S$), were carried out under natural weathering conditions. Direct photolysis of TB solution of concentration 4.3×10^{-3} M degraded 37.1% of the colored solution after two hours of solar irradiation, however, the solar photocatalytic decolorization percentage reached 79.04% and 86.21% after the addition of zinc oxide and titanium dioxide, respectively, for the same period. The degradation percentages were investigated by monitoring the dye decolorization spectrophotometrically. The decolorization rates of TB are markedly related with amount of hydroxyl radical formed. A suitable mechanism for the mineralization of TB has been proposed.

Keywords: Solar Photolysis and Photocatalytic decolorization, Thymol blue

Introduction

Arab countries are expected to face water shortages in the near future. Except Iraq and Syria, other Arab countries face serious water scarcity now, a problem which will be difficult¹ to be solved by 2025. Although Eritrea and Lebanon are regarded as water stress countries, other countries such as Algeria, Bahrain, Egypt, Jordan, Kuwait, Libya, Morocco, Oman, Palestine, Qatar, Saudi Arabia, Tunisia, United Arab Emirates and Yemen. are Classified as water scarce countries².

However, the availability of sunlight could recover this black image of water dearth. The direct solar radiation in Arab homeland is ranging from 4.1 kw h / m² /day in Mosul city in Iraq to 6.7 kw h / m² / day in Nouakchott, the capital of Mauritania^{3,4}. Moreover, sunshine duration ranges between 7.5 h in Tunisia to 10.7 h in Egypt³. Consequently, the water deficiency, which becomes worse due to the huge industrial consumption of water. Moreover, avoiding the recycling or reusing wastewater in other industries makes the problem more worse. However, investment of the available solar energy for the purpose of purification of industrial wastewater will lead to the dramatic reparation of the anticipated disastrous scarcity of the domestic water.

Decolorization is one of the most important nominated techniques utilized in the treatment of the industrial wastewater. Gizechulska and Morawski⁵ reported that the removal of color from wastes is more important than the removal of other colorless chemicals. However, not all the decolorization processes are suitable for decontamination. The degradation products should be non-toxic or at least less toxic than the degraded original compound⁶.

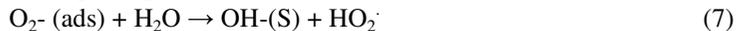
Recent studies demonstrated that heterogeneous photocatalysis is the most efficient technique in the degradation of colored chemicals⁷⁻¹⁵. These studies used titanium dioxide and / or zinc oxide in the photolysis processes. The band gap for both titanium dioxide and zinc oxide is (~ 3.2 eV), indicating that their photocatalytic activities are shown only under UV illumination⁸⁻¹⁰. Electrons in the valence band of semiconductors are excited to the conduction band when semiconductors are illuminated with UV radiation, leaving the positive hole in the valence band.

Photosensitization process occurs when a semiconductor mediates the oxidation of colored reagents⁶. In photosensitization, the adsorbed colored compound(s) on the surface of the semiconductor could absorb a radiation in the visible range¹¹⁻¹³. The excited colored dye (S*) (in the singlet or triplet state) will inject an electron to the conduction band of the semiconductor¹⁶.

Advanced oxidation processes (AOPs) concern with the total oxidation of organic and inorganic materials in wastewater by heterogeneous photocatalysis. The hydroxyl group radical (.OH); which is formed by the photocatalytic process, where the reaction is illuminated by UV in the presence of a semiconductor, or from the photosensitization processes; will oxidize all the organic to CO₂ and H₂O (mineralization). This is attributed to very strong oxidizing power of .OH which has the standard redox potential of + 2.8 V.



However, in the presence of a sensitizer, the following reactions will take place:



The TB, whose structure is shown in Figure 1, is used as biological stain and pH indicator. It was demonstrated that the compound can be harmful. Indeed, it is hazardous irritant and permeative substance. Inhalation: may cause irritation to the respiratory tract. Symptoms include coughing and shortness of breath.. Ingestion of large oral doses could cause irritation to the gastrointestinal tract. Skin Contact might cause irritation with redness and pain. Eye contact may cause irritation, redness and pain.

The present work, aims to study the photolysis and photocatalytic degradation of thymol blue using TiO₂ and ZnO under natural weathering conditions.

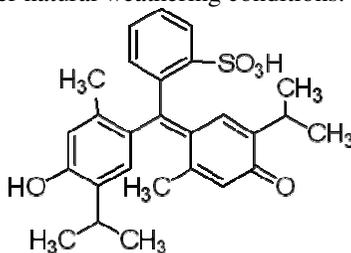


Figure 1. The structural formula of TB

Experimental

Materials and Methods

TB supplied by Fluka with a high purity (99%) was used as a test colored pollutant. Its structure is shown in Figure 1. It is usually used as pH indicator and as biological stain. Photocatalyst employed in this study was the titanium dioxide, supplied by BDH Limited with 98% purity (General purpose reagent GPR). Zinc oxide with 99% purity, supplied by BDH, is also utilized in this study as photocatalyst. All other chemicals were used without further purification.

TB was analyzed by UV-Visible Spectrophotometer (UV 2100 Unico). The determined wavelength was 565 nm, which is the maximum absorption wavelength. The determined absorption was conducted to a concentration through the Standard- Curve method of dye. The range of the concentration of aqueous propan-2-ol solution (1:1 percentage) was $4.0 \times 10^{-6} - 5.2 \times 10^{-6}$ M, which was used to conduct the calibration curve in this work.

Experiments were carried out during April and May 2005, over continuous 45 days period. All experiments have been performed at the floor of chemistry department building in Ibb University, in an open atmosphere between 11.00 a.m.-1.00 p.m. Sunlight illumination was accomplished in a 400 cm³ beaker containing 100 cm³ of the TB solution. The sunlight radiation was collected using converging lens with a focal length of 14 cm. In all experiments, 400 mg of the catalyst, anatase or zinc oxide, was suspended using a magnetic stirrer in dye solution. During the irradiation of the reaction mixture, 2 cm³ samples of the reaction mixture were, periodically, withdrawn using a syringe with a long pliable needle. These were centrifuged to separate the solid catalyst. In each case, 1 cm³ of the supernatant was drawn and diluted to 50 cm³ to measure the absorbance at 565 nm, using ultraviolet-visible spectrophotometer, type UV 2100, supplied by Unico Company. The photodegradation percentage of the dye was followed spectrophotometrically by a comparison of the absorbance, at specified interval times, with a calibration curve accomplished by measuring the absorbance, at known wavelengths, with different concentrations of the dye solution.

The experiments were also performed in the absence of the semiconductor in the presence and absence of sunlight. Decolorization efficiency is also calculated from a mathematical equation adapted from measurements of decolorization used before¹⁷⁻¹⁸.

$$\text{Decolorization} = \frac{(\text{Absorbance})_0 - (\text{Absorbance})_t}{(\text{Absorbance})_0} \times 100 \quad (9)$$

Where (Absorbance)₀ is the absorbance at 565 nm before irradiation and (Absorbance)_t is the absorbance at time t.

Results and Discussion

Figure 2 shows that TiO_2 gives good results compared with ZnO and in the absence of semiconductor (photolysis). However, no changes in color were observed in the absence of the catalyst and sunlight. The photolysis and photocatalytic degradation of TB were found to be fell in the sequence:

$$\text{Photocatalytic (TiO}_2\text{)} > \text{Photocatalytic (ZnO)} > \text{Photolysis}$$

The decolorization of TB in the presence of sunlight and TiO_2 or ZnO was increased linearly with irradiation time and was found to be equal to 86.21% in the presence of TiO_2 after 2 h of irradiation and 79.04% in the presence of ZnO for the same period.

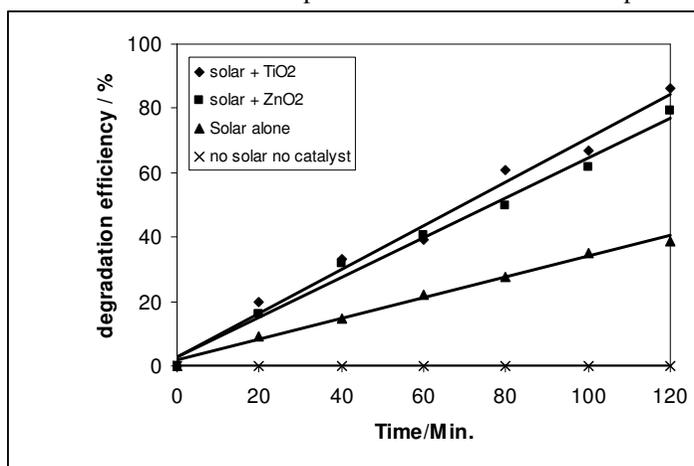


Figure 2. Degradation of TB under different conditions

No decolorization has been obtained in the absence of the sunlight and catalyst. However, in direct photolysis under sunlight process, *i.e.* in the absence of catalysts, the photolysis was observed to increase linearly with irradiation time to become 37.16% after 2h of irradiation. Comparison between the two methods used for the determination of decolorization efficiency is shown in Table 1.

Table 1. Comparison between decolorization efficiency obtained from the calibration curve and the adapted equation

Condition	Decolorization efficiency / %	
	Calibration Curve	Adapted equation
Solar + TiO_2	86.61	86.21
Solar+ ZnO	77.22	79.04
Solar alone	40.08	37.10
No solar, no catalyst	0	0

The degradation pattern of the TB could occur *via* the four main schemes shown in Figures 3-6, respectively, where the attack of the active hydroxyl radical ($\cdot\text{OH}$) upon the TB takes place via four suggested positions in the segments of TB mentioned in Figures 2-5. Figure 3 reveals that segment (I) is attacked by the $\cdot\text{OH}$ via the four main routes (a-d). The first route (a) includes the formation of the adsorbed intermediates (Ia), which are ultimately mineralized¹⁹ to CO_2 and H_2O . The second route (b) involves the formation of the dihydroxy intermediate (Ib), which is degraded¹⁹ to CO_2 and H_2O .

The third route (c) includes the attack of $\cdot\text{OH}$ on the isopropyl to form the isopropyl radical (Ic) with the evolution²⁰ of H_2O . The (Ic) intermediate is attacked by O_2 to give the corresponding peroxy intermediate (Id), which is degraded²⁰ to the alkene (Ij) derivative and HO_2 .

The fourth route (d) involves the transformation of the CH_3 group to the aldehydic group, CHO , to give the intermediate (If), which could be converted to other intermediate products, that are mineralized¹⁹ to CO_2 and H_2O . In addition, the aldehydic product (If) could be attacked by $\cdot\text{OH}$ to give the dihydroxy intermediate (Ig), which is degraded¹⁹ to CO_2 and H_2O . Moreover, the intermediate (If) may react with $\cdot\text{OH}$ with the elimination of H_2O to give the isopropyl radical derivative (Ih), which reacts with O_2 to yield the peroxy radical (Ii), which gives the alkene derivative²⁰ (Ij) beside $\cdot\text{HO}_2$.

Figure 4 demonstrates the degradation main two routes for segment (II). The first route (a) comprises the degradation²¹ of (II) into CO_2 , H_2O and H_2SO_4 . The second route (b) involves the introduction of $\cdot\text{OH}$ group to the benzene ring to form the intermediate (IIa), which is mineralized²¹ to CO_2 , H_2O and H_2SO_4 .

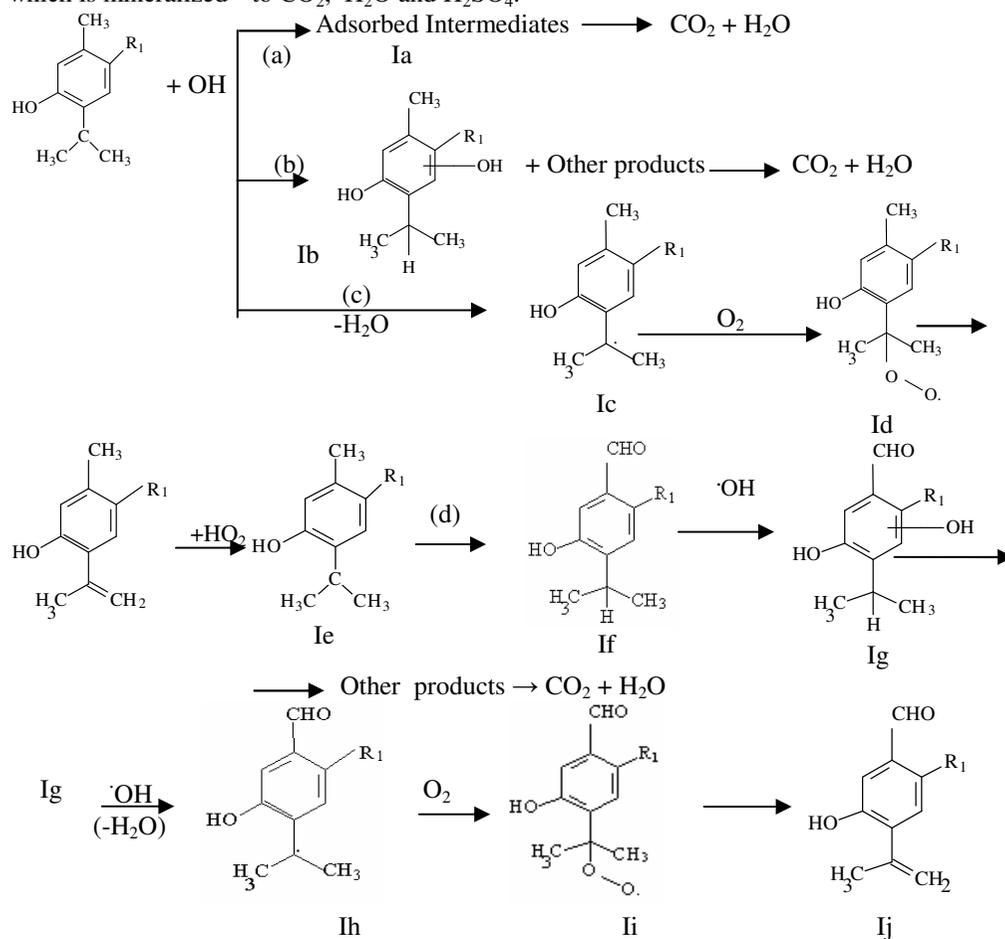


Figure 3. The degradation route of the segment (I).

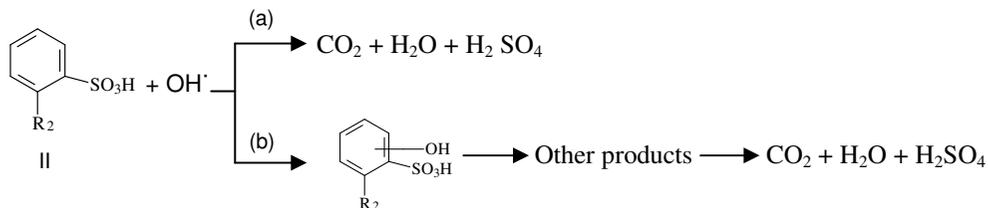


Figure 4. The degradation route of the segment (II).

Figure 5 reveals the main degradation route of the segment (III). Attack on the exo-double bond by $\cdot\text{OH}$ leads to the formation of alkyl radical (IIIa), which reacts with O_2 to give the peroxy radical (IIIb), which is converted to the tetroxide derivative (IIIc) that decompose to the radical (III d) as well as two molecules of the ketone (IIIe). Moreover, the product (III d) could be converted to the peroxy radical (III f) that is transformed to the benzoquinone intermediate (III g) besides HO_2 . It is noteworthy to mention that other routes could take place including the attack upon the isopropyl and the methyl groups to give other intermediates such as aldehydic, isopropyl, alkene intermediate which were outlined in the previous degradation routes.

Figure 6 shows the degradation pattern for the segment (IV), which occur through two routes (a and b). Route (a) includes the formation aldehydic derivative (IVa), via the attack upon CH_3 group by $\cdot\text{OH}$, which is degraded to other products¹⁹ that lead finally to CO_2 and H_2O . Route (b) involves the elimination of H_2O to give isopropyl radical derivative (IVb); which is oxidized by O_2 to yield the peroxy radical (IVc); that leads finally to the alkene derivative (IVd) and HO_2 .

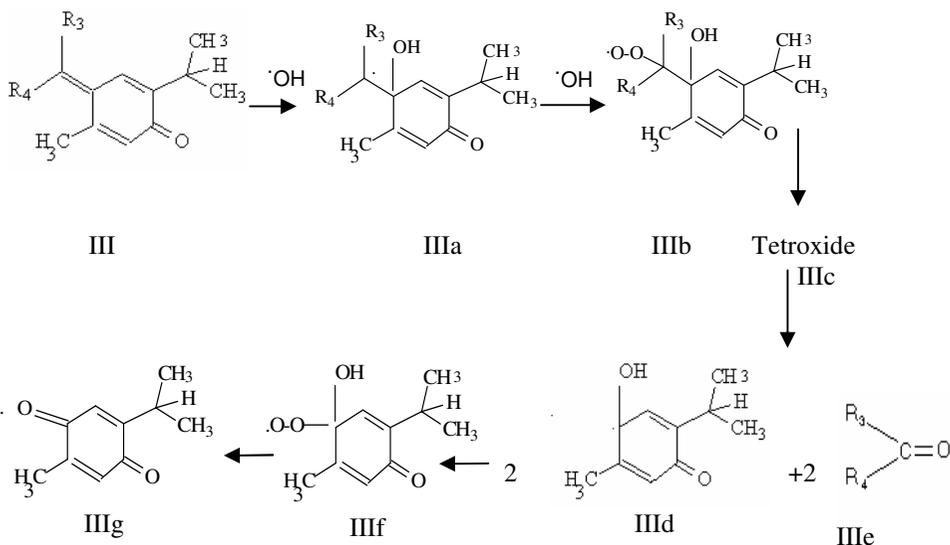


Figure 5. The degradation route of the segment (III).

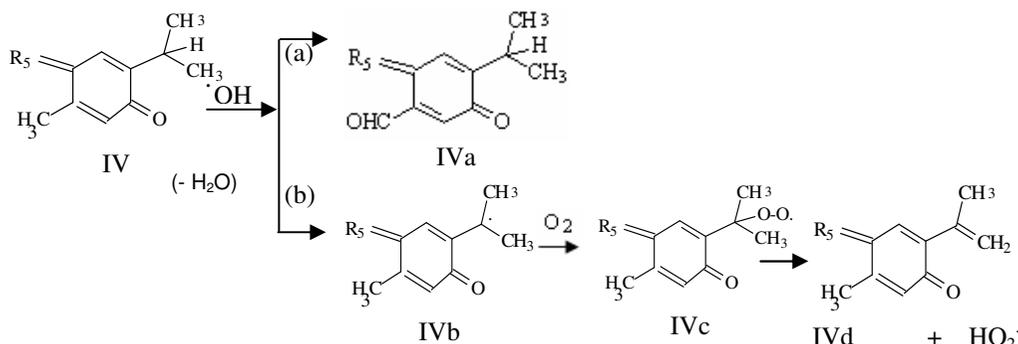


Figure 6. The degradation route of the segment (IV).

It is imperative to mention that the kinetics of the aromatic ring degradation were monitored at the higher wavelength²².

Conclusions

1. Solar energy is essential for photodegradation process of TB. The degradation percent reached 37.1% after 2 h of solar irradiation.
2. The percent of photodegradation is more than doubled in the existence of the photocatalyst.
3. The band gap of TiO_2 and ZnO is equal to ~ 3.2 eV, *i.e.*, near ultraviolet irradiation is needed to excite the electrons from the band gap of the semiconductor to the conduction band. However, the existence of the colored organic compound enables the utilization of the solar irradiation by photosensitization process, where the colored compound could be excited by visible light and then injects an electron to the conduction band of the photocatalyst.
4. The availability of the solar energy in the Arab region can be vastly invested in the treatment of the industrial wastewater, especially the colored one. Among the different wastewater treatment techniques such as ozonation, biological treatment, Fenton's oxidation, electrochemical destruction and photosensitization, the latter was nominated to be the most preferable means.

References

1. Gardner-Outlaw and Engelman, Sustaining Water, easing Scarcity:: A Second update, Washington, D. C., Population Action International, 1997.
2. World Resources 1992-1993, Pacific Institute for Studies in: Development Environmental and Security Stockholm Environment Institute and World Bank Estimates, 1995.
3. Alnaser W E, Trieb F and Knies G, Possible cooperation between Arab and European Countries in Energy, *Water and Environmental Issues*, Sirse Forum, 2004 pp 99-100.
4. ESCWA Energy Options for Water Desalination in selected ESCWA members Countries, United Nations, New York, 2001.
5. Grzechulska J and Morawski A, *Appl. Catal. B: Environmental*, 2002 **36** (1), 45.
6. Jhon A, Jennifer T, Ken A, Heather R and Mathew T, *J. Chem. Edu.* 1999, **76** (12), 1680 [and references therein].

7. Hussein F, Mashkoor M and Al-Sharafy A, *National J. Chemistry*, 2003, **9**: 94.
8. Li X, Liu H., Cheng L and. Tong H, *Environ. Sc. Technol.*, 2003 , **37 (17)**, 3989.
9. Vione D, Picatonito T and Carlotti M E, *J. Cosmet Sci.*, 2003, **54**, 513.
10. Antharjanam S, Philip R and. Suresh D, *Ann. Chim.*, 2003 , **93 (9-10)** , 719.
11. Fernandez-Ibanez P, Planko J, Maitato S and de las Nieres F, *Water Res.*, 2003, **37 (13)** , 3180.
12. Ohno T, *Water Sci. Technol.* 2004, **49 (4)**, 159.
13. Alkhateeb A, Hussein F and Asker K, *Asian J. Chem.*, 2005, **17 (2)**, 1155.
14. Liu HL-Zhou D, Li X Z and Yue P T, *J. Environ. Sci. (China)*, 2003, **15 (5)**, 595.
15. Chen J, Liu M, Zhang J, Ying X and Jin L, *J. Environ. Manage*, 2004, **70 (1)**, 43.
16. Hussein F H and. Alkhateeb A N, *Desalination* 2007, **209**, 361.[and references therein].
17. Hongre D and Alkesson G, *Water Res.*, 1996, **30**, 2771.
18. Anaral P, Fernades D, Tarares A, Xarares A, Cammarota H, Continho J and Coelho M, *Environ. Technol.* 2004 , **25 (11)** ,1313.
19. Addamo M, Augugliaro V, Coluccia S, Di Paola A, Garc'ia-Lopez E., Loddo V, GianmarioMartra G M and Palmisano L, *Int. J. Photoenergy*, 2006, 1.
20. Maria Hoeben W F, Ph.D thesis, Eindhoven: Technische Universiteit Eindhoven, 2000.
21. Bahnemann D, Cunnigham J, Fox M A, Pelizzetti E, Pichal P and Sorpone N, *Aquatic and Surface Photochemistry*, Lewis Publishers, London, 1994, p 261.
22. Topalov A, Sojic D, Molnar-Gabor D, Abramovic B and Comor M, *Applied Catalysis B: Environmental* , 2004, **54**, 125.



Hindawi

Submit your manuscripts at
<http://www.hindawi.com>

