Figure S1. The electronic absorption of aqueous solution of **1** upon addition of (a) Zn^{2+}, (b) Fe^{3+}, (c) Ni^{2+}, (d) Co^{2+}, (e) Na^{+}, and (f), K^{+}

Figure S2. The electronic absorption of aqueous solution of **1** (5.0×10^{-6} M) upon addition of (a) Hg^{2+} (0-10 equivalent), (b) Pb^{2+} (0-10 equivalent), (a) Cd^{2+} (0-10 equivalent), and (a) Cu^{2+} (0-10 equivalent) ions

Figure S3. The electronic absorption spectra of aqueous solution of **1**(5.0×10^{-6} M) upon addition of increasing amount [0.12, 0.25, 0.38, 0.51, 0.63, 0.89,and 1.27 ($\times 10^{-6}$ M)] of Hg^{2+}, Pb^{2+},Cd^{2+}, and Cu^{2+} ions. The inset displays a zoomed view of the visible portion of the spectra

Figure S4. The electronic absorption of aqueous solution of **1** (5.0×10^{-6} M) upon addition of increasing amount (a) [0, 0.5, 1.25, 2, 2.5, and 3.75 ($\times 10^{-6}$ M)] Cu^{2+} and Hg^{2+} (b) [0, 0.5, 1.25, 2.5, 3.75, and 5 ($\times 10^{-6}$ M)] Cu^{2+} and Pb^{2+} (c) [0, 0.5 , and 1.25 ($\times 10^{-6}$ M)] Cu^{2+} and Cd^{2+} ions

Figure S5. The electronic absorption of aqueous solution of **1** (5.0×10^{-6} M) upon addition of Ca^{2+} (0-16 equivalent)

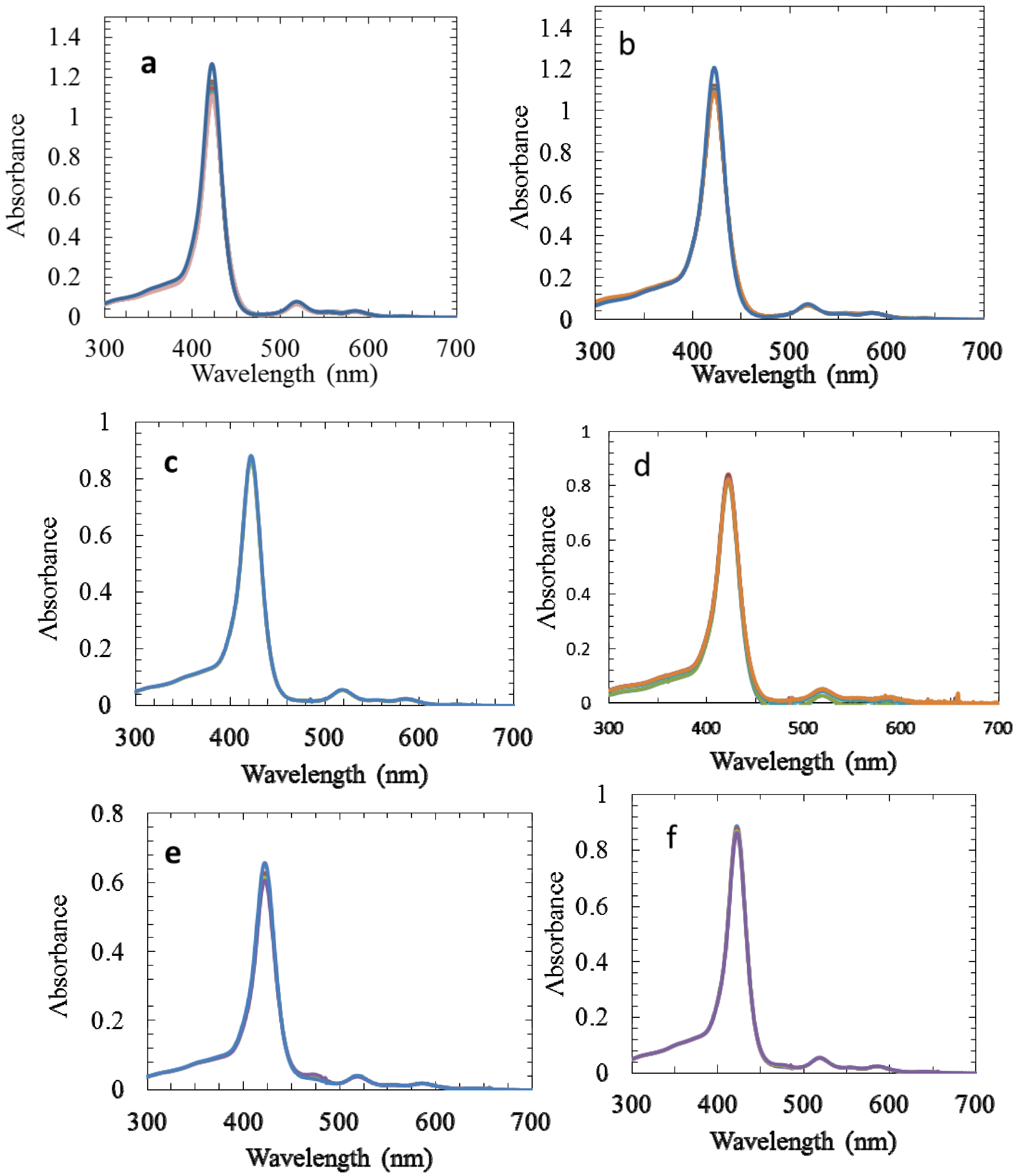

Figure S1. The electronic absorption of aqueous solution of **1** upon addition of (a) Zn^{2+}, (b) Fe^{3+}, (c) Ni^{2+}, (d) Co^{2+}, (e) Na^{+}, and (f), K^{+}

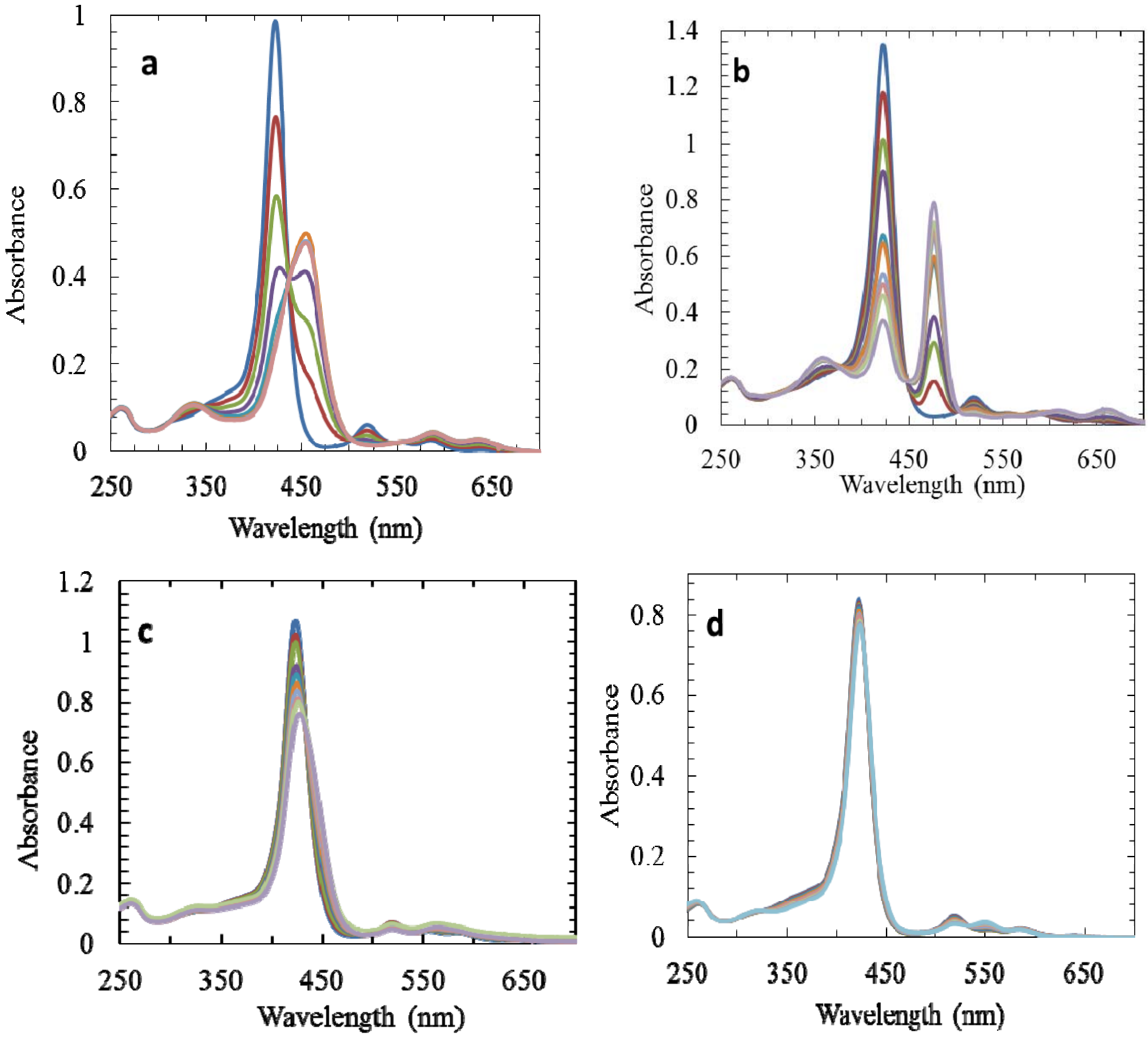

Figure S2. The electronic absorption of aqueous solution of **1** (5.0×10^{-6} M) upon addition of (a) Hg^{2+} (0-10 equivalent), (b) Pb^{2+} (0-10 equivalent), (a) Cd^{2+} (0-10 equivalent), and (a) Cu^{2+} (0-10 equivalent) ions

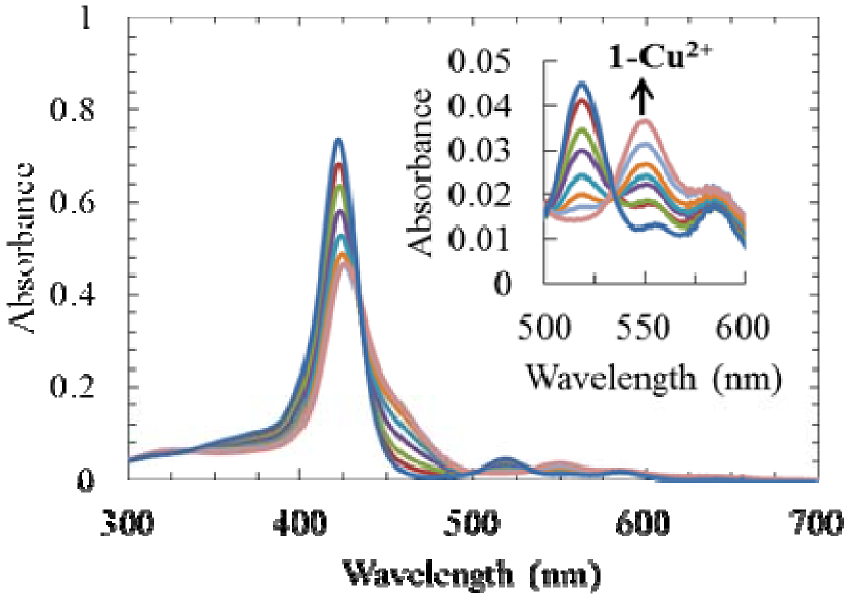

Figure S3. The electronic absorption spectra of aqueous solution of **1**(5.0×10^{-6} M) upon addition of increasing amount [0.12, 0.25, 0.38, 0.51, 0.63, 0.89,and 1.27 ($\times 10^{-6}$ M)] of Hg^{2+}, Pb^{2+},Cd^{2+}, and Cu^{2+} ions. The inset displays a zoomed view of the visible portion of the spectra

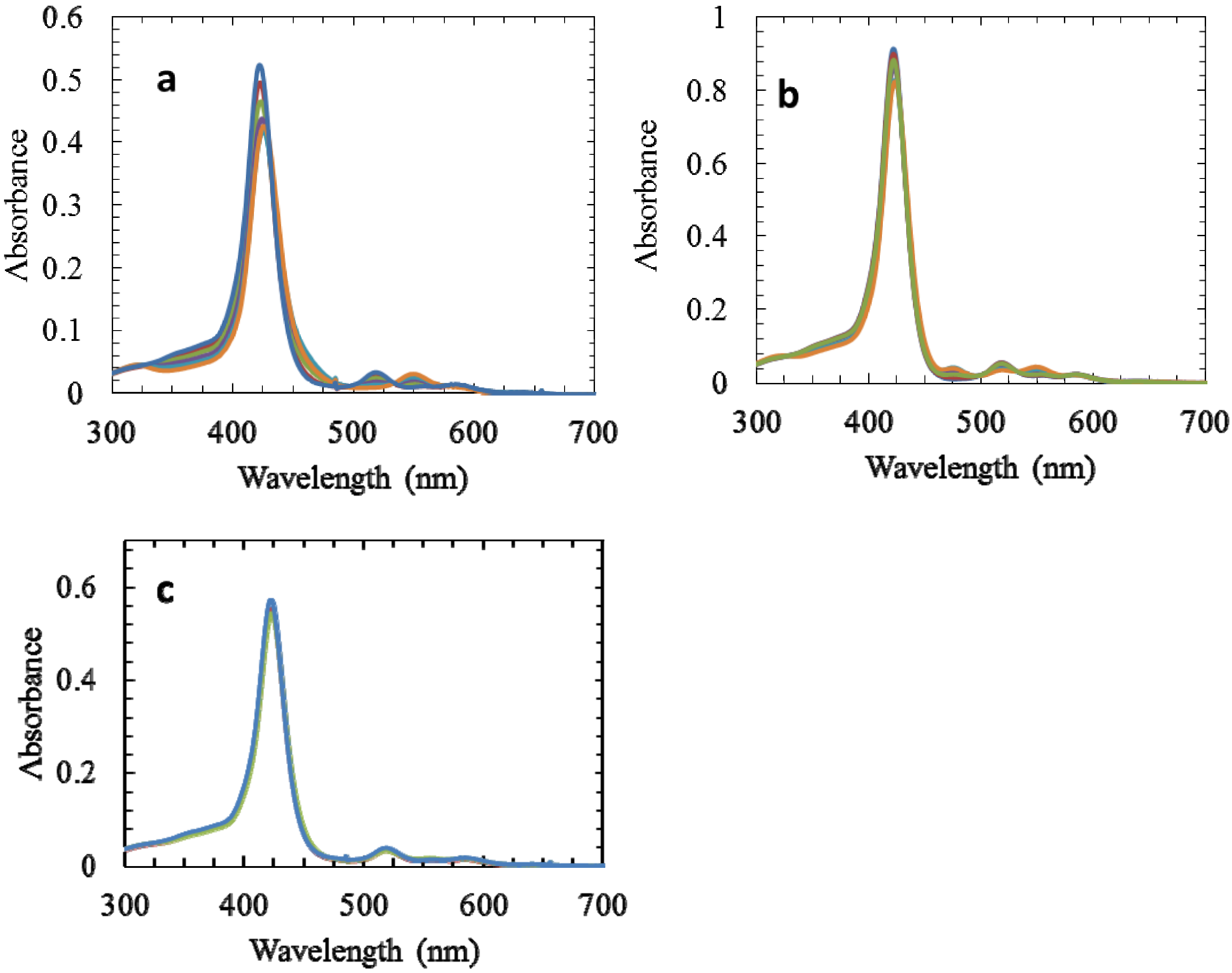

Figure S4. The electronic absorption of aqueous solution of **1** (5.0×10^{-6} M) upon addition of increasing amount (a) [0, 0.5, 1.25, 2, 2.5, and 3.75 ($\times 10^{-6}$ M)] Cu^{2+} and Hg^{2+} (b) [0, 0.5, 1.25, 2.5, 3.75, and 5 ($\times 10^{-6}$ M)] Cu^{2+} and Pb^{2+} (c) [0, 0.5 , and 1.25 ($\times 10^{-6}$ M)] Cu^{2+} and Cd^{2+} ions

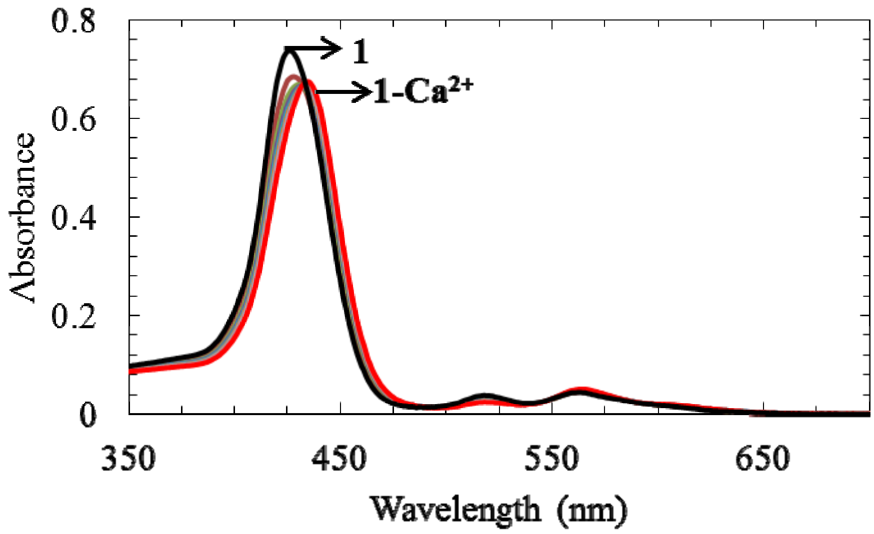

Figure S5. The electronic absorption of aqueous solution of **1** (5.0×10^{-6} M) upon addition of Ca^{2+} (0-16 μL of 1×10^{-3} M)