Supplementary Material

Design and Synthesis of a Dinuclear Copper(II) Probe for Selective Fluorescence Sensing of Pyrophosphate

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Figure S1. ¹H NMR of compound **DPAC** in CDCl₃.







Figure S4. UV-vis absorption spectra of compound DPAC (10 µM) upon addition of various metal

Wavelength (nm)

ions (2 equivalents) in CH₃CN: HEPES (3: 2, v/v, pH=7.2) solutions.



Figure S5. The Benesi-Hildebrand fitting of titration plots with the titration of Cu^{2+} assuming 1:2

stoicheiometry with Cu²⁺.



Figure S6. Emission responses of compound DPAC (10 µM) upon additions of various metal ions

in CH₃CN: HEPES (3:2, v/v, pH=7.2) solution.



Figure S7. ESI-MS spectra of **DPAC** in the presence of Cu²⁺ in CH₃CN solution.



Figure S8. The simulated (left) and experimental (right) isotopic distributions of



 $[Cu_2(DPAC)+ClO_4]^{3+}$ species.



 $[Cu_2(DPAC)+2ClO_4]^{2+}$ species.



Figure S10. Fluorescence intensity of DPAC-Cu²⁺ (10 µM) in CH₃CN: HEPES (3:2, v/v, pH=7.2)

solution upon additions of PPi in range of 1-4 μ M.



Figure S11. ESI-MS spectrum of **DPAC**-Cu²⁺ with the addition of 1 equivalent of PPi.





 $[Cu_2(DPAC)+PPi+H]^+$ species.



Figure S13. HPLC analyses of **DPAC**; **DPAC**+2 equiv.Cu²⁺; **DPAC**+2 equiv.Cu²⁺+0.5 equiv. PPi; **DPAC**+2 equiv.Cu²⁺+1 equiv. PPi. Liquid phase conditions: MeOH-water (8:2, v/v, mobile phase); 1.0 mL min⁻¹ (current velocity); ZORBAX SB-C18 (chromatographic column).



Figure S14. Fluorescence spectra of compound **DPAC** (10 μ M) in different buffer solutions.